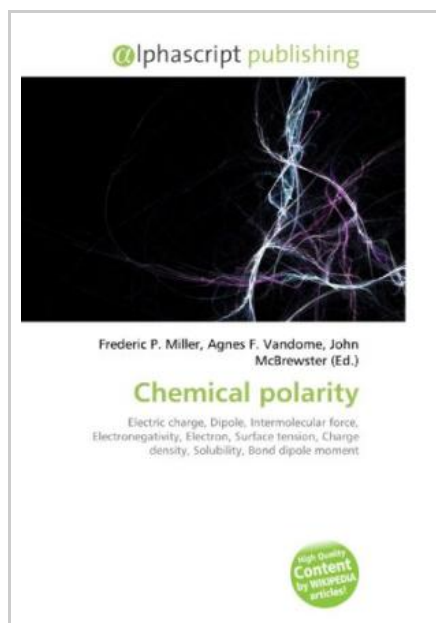




Chemical polarity

By Frederic P. Miller

Alphascript Publishing Dez 2009, 2009. Taschenbuch. Book Condition: Neu. 220x150x7 mm. Neuware - High Quality Content by WIKIPEDIA articles! In chemistry, polarity refers to a separation of electric charge leading to a molecule having an electric dipole. Polar molecules can bond together due to dipole. Dipole intermolecular forces between one molecule (or part of a large molecule) with asymmetrical charge distribution and another molecule also with asymmetrical charge distribution. Molecular polarity is dependent on the difference in electronegativity between atoms in a compound and the asymmetry of the compound's structure. For example, a molecule of water is polar because of the unequal sharing of its electrons in a 'bent' structure, whereas methane is considered non-polar because the carbon shares the electrons with the hydrogen atoms uniformly. Polarity underlies a number of physical properties including surface tension, solubility, and melting and boiling-points. 120 pp. Englisch.



READ ONLINE

[7.32 MB]

Reviews

This publication is amazing. It is definitely basic but shocks in the fifty percent of your publication. You wont feel monotony at anytime of your own time (that's what catalogues are for concerning if you question me).

-- **Prof. Kirk Cruickshank DDS**

This kind of book is every little thing and taught me to looking ahead of time and a lot more. I am quite late in start reading this one, but better then never. I found out this book from my dad and i encouraged this pdf to find out.

-- **Justus Hettinger**